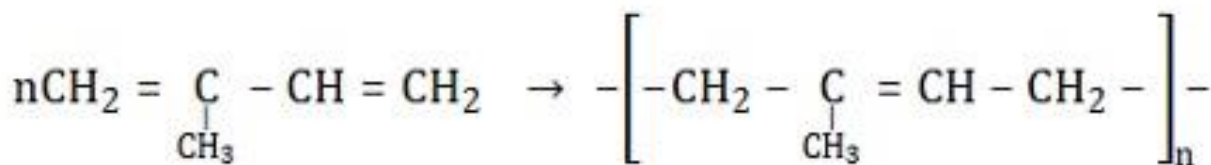
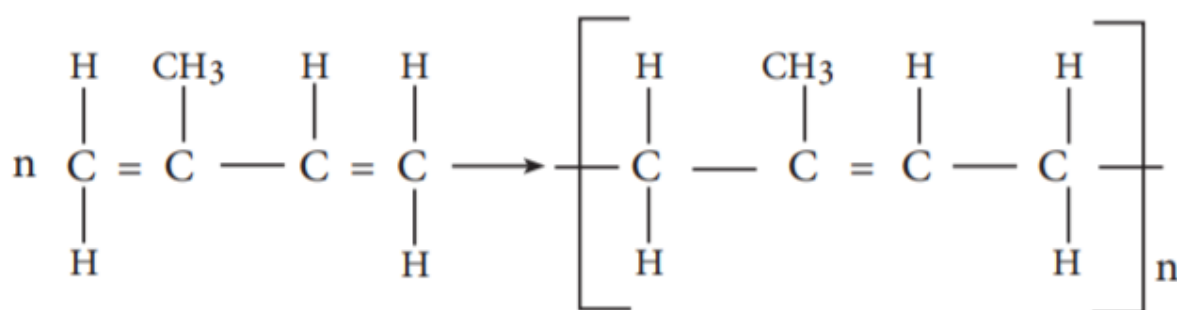
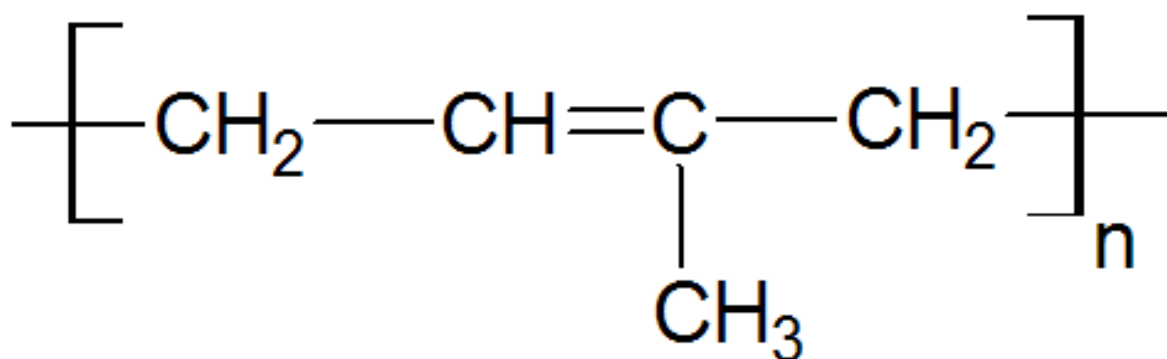


- 4) bond of the bond between the and the C atoms will break.
- 5) Each is made up of 2
- 6) One will move of the molecule and the other will move to the position between the and the C atoms.
- 7) Therefore now there will be a bond between the and the C atoms.
- 8) One of the bond between the and the C atoms will break.
- 9) Each is made up of 2
- 10) One will move the molecule and the other will move to the position between and C atoms.
- 11) Therefore now there will be a bond between the and the C atoms.
- 12) The 2 which came to the position between the and C atoms will get together and form a new
- 13) Therefore the new and the earlier between the and the C atoms will now become a bond.
- 14) There will be an at the of the molecule and at the of the molecule.
- 15) The behind of the molecule and the electron of the molecule will get together and form a new

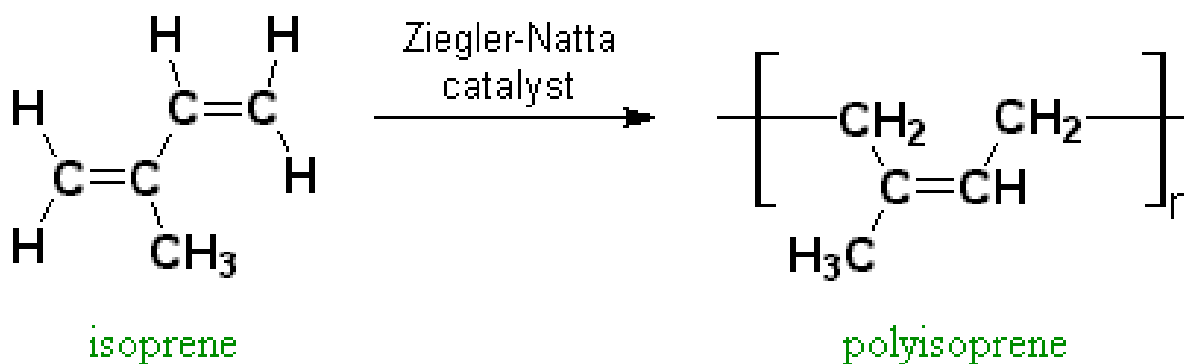
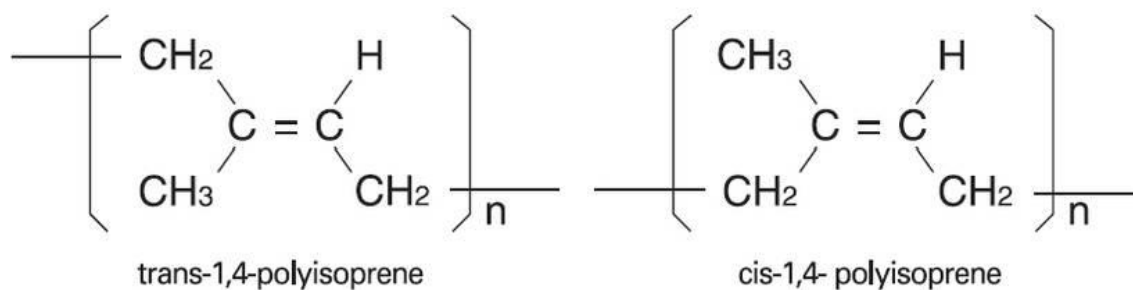
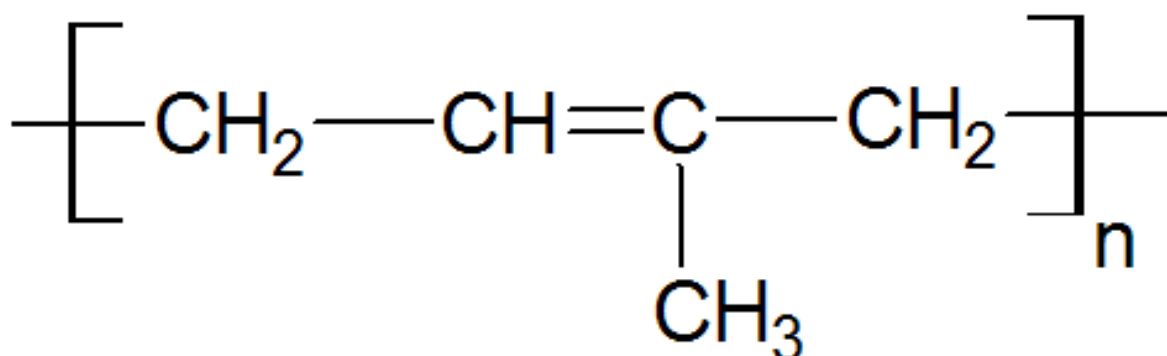


Isoprene

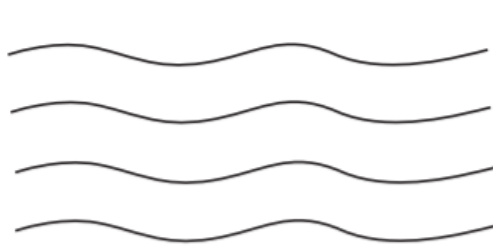
[2 - methyl butadiene]

Natural rubber

[Poly isoprene]



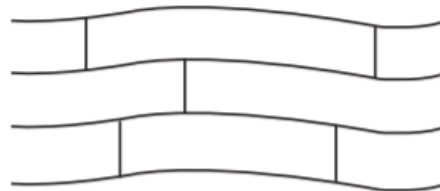
Types of polymers



Linear polymers
Figure 14.5



Polymers with side chains
Figure 14.6

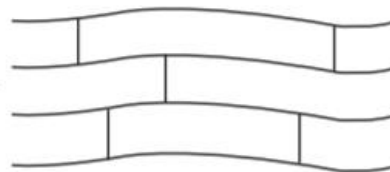


Polymers with cross links

Vulcanised rubber



Rubber



Vulcanized rubber

- 1) is heated with and produced rubber.
- 2) Properties of rubber
 - (i) Less and
 - (ii) Good of electricity
 - (iii) Can make various
- 3) When rubber is it will give out

(i) CO ₂	(ii) CO	(iii) H ₂ O	(iv) SO ₂
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Why are popular

- 1)
- 2) available
- 3) Ability to synthesis with required
- 4) Easy to use
- 5) Ability to make different
- 6) Ability to synthesis with required

Disadvantages of artificial polymers

- 1) Not Therefore not through processes. Therefore get accumulated in the
- 2) Combustion of produce gases.
- 3) Cloths produced using & do not absorb Therefore cause discomfort to body. (This can be reduced by mixing them with polymers such, & with artificial polymers such as &)