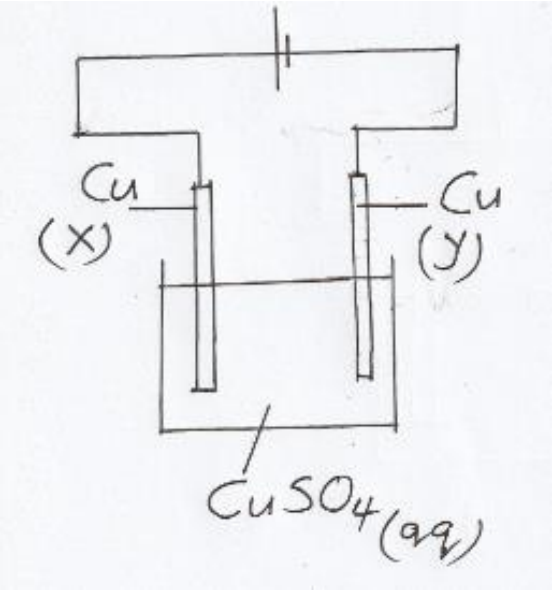
Grade	Institute	Day	Time	Starting Date
10	Channa Asela Institute, Mt. Lavinia	Saturday	02.30pm – 05.30pm	5th September
10	Shakthi Institute, Colombo 04	Friday	02.30pm – 04.30pm	4th September
10	Shakthi Institute, Waterfront Colombo 2	Tuesday	02.30am – 04.30pm	8th September
11	Channa Asela Institute, Mt. Lavinia	Saturday	10.30am – 01.30pm	5th September
11	Shakthi Institute, Colombo 04	Monday	02.30pm – 04.30pm	7th September
11	Shakthi Institute, Waterfront Colombo 2	Wednesday	02.30pm – 04.30pm	2nd September
Past Paper	Channa Asela Institute, Mt. Lavinia	Saturday	06.00pm – 08.00pm	5th September
Past Paper	Shakthi Institute, Colombo 04	Monday	04.45pm – 6.45pm	7th September
Past Paper	Shakthi Institute, Waterfront Colombo 2	Wednesday	0.4.45pm – 06.45pm	2nd September

Online Class	Days	Time
Grade 10	Mondays & Fridays	8.30pm – 9.45pm
Grade 11	Tuesdays & Thursdays	8.30pm – 9.45pm
Grade 10 Repeat	Saturdays	8.30pm – 9.45pm
Grade 11 Repeat	Sundays	8.30pm – 9.45pm

Electrolysis of aqueous $CuSO_4$ solution using Cu



1) The electrode (Y) which is connected to the …… terminal of the battery is the …… electrode.

2) The electrode (X) which is connected go the …….. terminal of the battery is the …… electrode.

3) Aqueous $CuSO_4$ contains …….. and ……..

4) $CuSO_4$ dissociates into ……. and ……

$CuSO_4 \rightarrow$ ……. + …….

5) H_2O dissociates into ……. and …….

$H_2O \rightarrow$ ……. + …….

6) ……. move out from the ……. terminal of the battery to the …… electrode (….).

7) The ……. which came to the ……. electrode (…….) will be given to the ……..

8) …….. are ……. charged.

9) Therefore the …… charged …… will be taken up by ……. charged ions (……).

10) The ……. ions (……) found in the solution are …… and ……

11) The ……. ions (…….) of the ……. reactive metal will …… electrons and become ……

12) ……. is less reactive than …….

13) Therefore …… ions will accept two …… charged ……. and become …….

……. + …… → ……..

14) The reactions which …… electrons are called ……. reaction

15) Therefore the …… reaction is ……. + ……. → ……..

16) The place where ……. reactions occur is called the ……

17) Therefore ……. electrode will be the ……

18) The reaction which occur near the …… is called ……. reaction

19) The cathodic reaction is ……. + ……. → …….

20) Since …… is a metal, …… will …… electrons and send them to the ……. terminal of the battery. Eg. …… → …….. + ……..

21) The reactions which ……. electrons are called …… reactions.

22) Therefore the ……. reaction is …… → ……. + …….

23) The place where …… reactions happen is called the ……..

24) Therefore ……. electrode will be the …….

24) The reaction which happens near the …… is called the ……. reaction.

24) Therefore the …… reaction is …… → …… + …….

25) The remaining ……. and ……. ions will now get together and form …….

…….. + …….. → ……

Questions

1) Negative electrode = …….

2) Positive electrode = …….

3) Anode = …….

4) Cathode = ……

5) Anodic reaction = …….

6) Cathodic reaction = …….

7) Electrons moved = ……

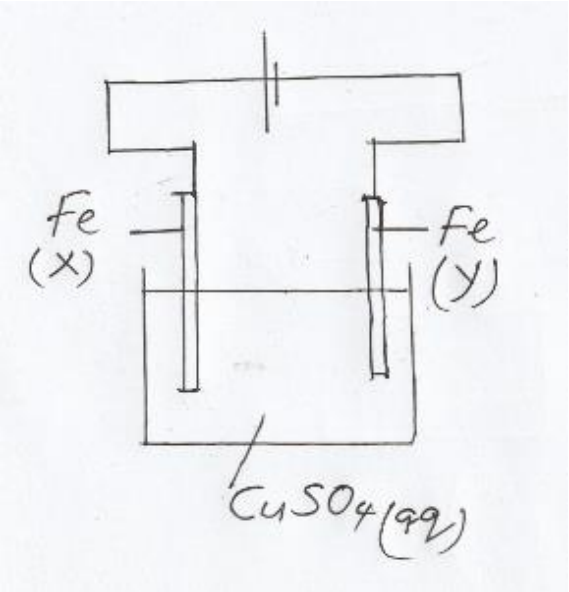
8) Current moved = …….

10) Observations

(i) X electrode will …….

(ii) A reddish brown substance will deposit on ……. electrode

Electrolysis of aqueous $CuSO_4$ solution using Fe electrodes



1) The electrode (Y) which is connected to the ……. terminal of the battery is the ……. electrode.

2) The electrode (X) which is connected go the ……. terminal of the battery is the ……. electrode.

3) Aqueous $CuSO_4$ contains ……. and …….

4) $CuSO_4$ dissociates into …… and …….

$CuSO_4 \rightarrow$ ……. + …….

5) H_2O dissociates into ……. and …….

$H_2O \rightarrow$ ……. + …….

6) …… move out from the ……. terminal of the battery to the …… electrode (……..).

7) The ……. which came to the …… electrode (…….) will be given to the solution.

8) …… are ……. charged.

9) Therefore the …… charged …….. will be taken up by ……. charged ions (……).

10) The …….ions (+…….) found in the solution are …….. and ……

11) The ……. Ions (……..) of the …….. reactive metal will ……. electrons and become …….

12) …… is less reactive than ……

13) Therefore …….. ions will …….. two …… charged …… and become …….

…….. + ……. → ……

14) The reactions which ……. electrons are called …… reaction

15) Therefore the ……. reaction is …….. + …… → ……

16) The place where …… reactions occur is called the ……..

17) Therefore …… electrode will be the ……..

18) The reaction which occur near the …… is called ……. reaction

19) The …….. reaction is …….. + ……. → ……

20) Since …… is a metal ……. will …… electrons and send them to the ……. terminal of the battery. Eg. …….. → …… + …….

21) The reactions which ……. electrons are called …….. reactions.

22) Therefore the ……. reaction is ……. → …….. + ……..

23) The place where ……. reactions happen is called the ……..

24) Therefore …….. electrode will be the ……..

24) The reaction which happens near the …….. is called the ……. reaction.

24) Therefore the …… reaction is …… → …….. + …….

25) The remaining ……. and …… will get together and form …….

……. + ……. → …….

Questions

1) Negative electrode = ……

2) Positive electrode = ……

3) Anode = …….

4) Cathode = ……

5) Anodic reaction = …….

6) Cathodic reaction = ……..

7) Electrons moved = …….

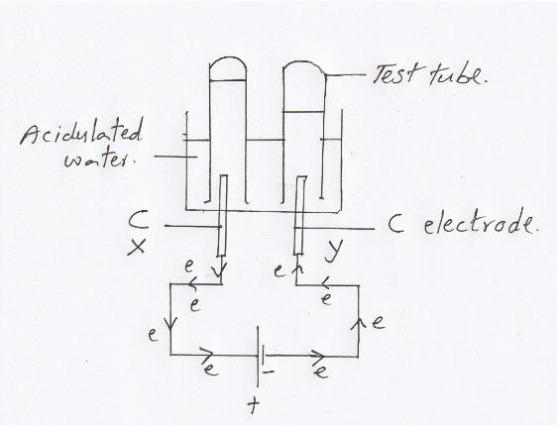
8) Current moved = ……..

10) Observations

(i) X electrode will …….

(ii) A reddish brown substance will deposit on ……. electrode

Electrolysis of acidulated water



1) Water is electrolyzed in a ……..

2) Acidulated water contains …….. and ……

3) H_2SO_4 dissociate into …… and ……

……. → ……. + …….

4) …… dissociates into …….. and ……

…… → ……. + ……..

5) The C electrode which is connected to the ……. terminal of the battery will be the ……. electrode.

6) Therefore …… electrode will be the …….. electrode.

7) The C electrode which is connected to the ……. terminal of the battery will be the …….. electrode.

6) Therefore …… electrode will be the ……. electrode.

7) ……. charged ……. move out from the …….. terminal of the battery and flow to ……. electrode.

8) The ……. charged …… which came to the …… electrode will be given to the ……..

9) Electrons are ……. charged.

10) The ……. charged …….. will be taken up by a ……. charged ion.

11) There is only one type of ……. charged ion. Eg – ……

12) These ……. charged …….. will move to the …….. electrode and ……. negatively charged ……. …… + ….. → …….

13) Adding ……. are called ……. reactions.

14) Therefore ……. + ……. → ……. is the ……. reaction.

15) The place where …… reactions occur is called the …….

16) Therefore the …….. electrode is the ……

17) The reaction which happens near the …… is called the …… reaction.

18) Therefore the …… reaction is …… + …… → ……

19) The …….. gas produced near the ……. (the …….. electrode) will displace down some …….. in the …… test tube and get collected at the ……. of the …… test tube.

20) Like the way …… moved from the ……. terminal of the battery to the ……. electrode, the ……will flow from the ……. electrode to the …… terminal of the battery.

21) There are 2 ……. charged ions in the …….. (acidulated water). Those are …… and …….

22) ……. will move to the …….. electrode and will ……. electrons and produce …….and …….

…… → ……. + …….. + ……

23) The reactions which ……. electrons are called ……. reactions.

24) Therefore …… → ……. + ……. + …….. reaction is the ……. reaction.

25) The place where the ……. reaction happens is called the ……

26) Therefore the …… electrode is the …….

27) The reaction which happens near the …… is called the …… reaction.

28) Therefore ……. → …… + ……. + …… is the ……. reaction.

29) The ……. gas which was produced near the ……. electrode will displace some water in the …… test tube and get collected at the top of the ……. test tube.

30) …… flows in the …… direction to the flow of ……

31) Therefore ……. flows from the …… terminal of the battery to the …….. electrode and from the …… electrode to the …… terminal of the battery.

32) The chemical formular of …… is ……..

33) Therefore the ratio of H:O in a …… molecule is ……

34) Therefore during the …….. of ……, the volume ratio of H_2 is to O_2 will be ……

Questions

1) Name the equipment use to electrolyze water : ……

2) Why few drops of acids are added to water : …….

3) Name the (+) electrode : …….

4) Name the (-) electrode : ……

5) Name the anode : …….

6) Name the cathode : …….

7) Write the anodic reaction : ……

8) Write the cathodic reaction : ……

9) The gas collected near the X electrode : ……..

10) The gas collected near the Y electrode : ……..

11) The volume ration of O_2 : H_2 : ……