0714 820 596 **Channa Asela** **www.OLscience.com**
Teacher, S. Thomas' College, Mt. Lavinia

Grade	Institute	Day	Time	Starting Date
10	Channa Asela Institute, Mt. Lavinia	Saturday	02.30pm – 05.30pm	5th September
10	Shakthi Institute, Colombo 04	Friday	04.30pm – 06.30pm	4th September
11	Channa Asela Institute, Mt. Lavinia	Saturday	10.30am – 01.30pm	5th September
11	Shakthi Institute, Colombo 04	Sunday	10.00am – 12.00pm	6th September
Past Paper	Channa Asela Institute, Mt. Lavinia	Saturday	06.30pm – 08.30pm	5th September
Past Paper	Shakthi Institute, Colombo 04	Sunday	01.00pm – 3.00pm	6th September

Online Class	Days	Time
Grade 10	Mondays & Fridays	8.30pm – 9.45pm
Grade 11	Tuesdays & Thursdays	8.30pm – 9.45pm
Grade 10 Repeat	Saturdays	8.30pm – 9.45pm
Grade 11 Repeat	Sundays	8.30pm – 9.45pm

Chapter 12 – Electrochemistry

Reactivity series

1) …… are arranged according to the …… order of their ……..
2) This series of …….. is called the ……..
3) K, ……., Ca, ……., Al, ……., Sn, ……., H, ……., Hg, ……., Pt, …….
4) Therefore, the …… which are on the ……. side of the reactivity series are …… reactive than the ……. which are on the …… side of the reactivity series.

   Eg. Mg is …… reactive than Al

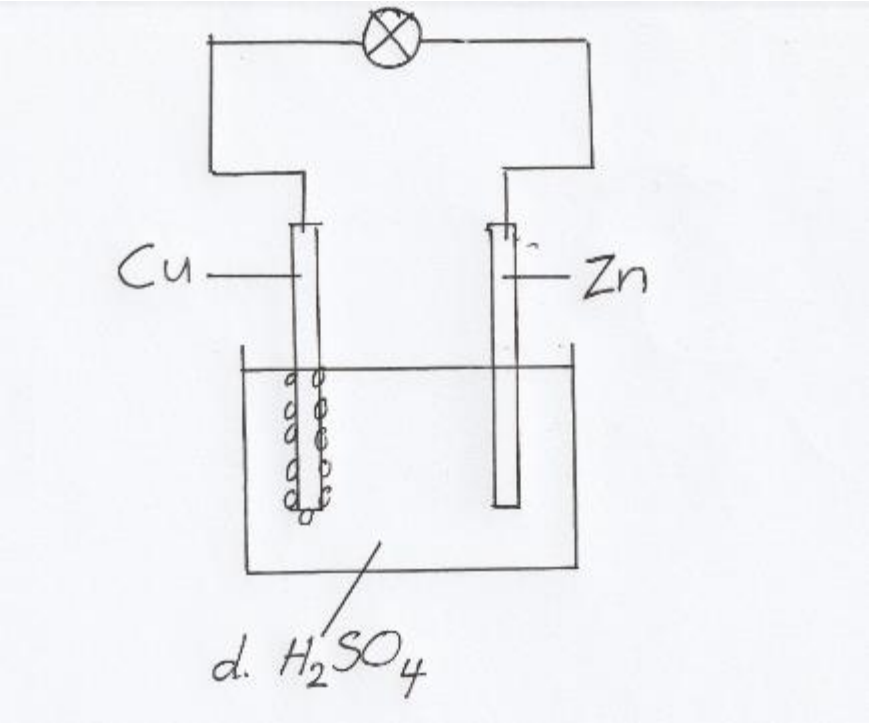
5) The neutral …….. of the …… reactive …….. will …… electrons according to its ……. and become …….. (…….. ions).

   Eg. K → …….. + ……..
   Mg → ……. + ……..
   Al → ……. + ……..

6) The ……. (…… ion) of the ……. reactive ……. will ……. electrons according to its valence and become a neutral ……

   Eg. H^+ + ……. → ……..
   $2H^+$ + ……. → ……..
   Ag^+ + ……. → ……

Electrochemical cell



Diluted …… acid was added to a …….. A …… metal and a …….. metal were dipped in diluted …….. acid. The two …….. were connected through a ……..

1) Diluted …….. acid contains …… and ……..
2) ……. dissociates into …….. and ……

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4) …… is more reactive than ……..

5) The valence of …… is ……..

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…….. → …… + ……..

7) Therefore the …… metal will …… slowly and the …….. ions will enter the ……..

8) The reactions which ……. electrons are called ……. reactions.

9) Therefore …….. → …… + ……. is an …….. reaction.

10) The place where ……. reactions occur is called the ……..

11) Therefore the …… metal is the ……..

12) The reaction which take place near an ……. is called an …… reaction.

13) Therefore …….. → …….. + ……. is an …….. reaction.

14) The electrons which were …… by the ……. metal will flow through the wire from the ……. metal to the ……. metal.

15) The electrons which came through the wire from the …… metal to …… metal will now enter the ……..

16) Electrons are …….. charged.

17) Therefore the …… charged electrons will be ……. by a …….. charged ion.

18) There are two …….. charged ions in the ……

Eg. …… and ……

19) The ions of the ……. reactive metal will …… the electron(s) according to its …….. and become ……..

20) ……. is less reactive than ……. (please check the reactivity series)

21) Therefore two ……. ions will ……. two electrons and become ……. gas.

…… + …….. → ……

22) Therefore ……. bubbles will emit from the ……. metal.

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29) ……. flows in the …… direction to the flow of electrons.

30) Therefore the …….. will flow from …… to ……

31) ……. flows from …… to ……..

32) Therefore the …… metal is the …….. electrode and the …… metal is the ……. electrode.

Questions

1) Name the cathode ……

2) Name the anode ……..

3) Write the cathodic reaction ……..

4) Write the anodic reaction ……

5) Write the direction of the electron flow ……

6) Write the direction of the current flow ……

7) Name the positive electrode ……..

8) Name the negative electrode ……

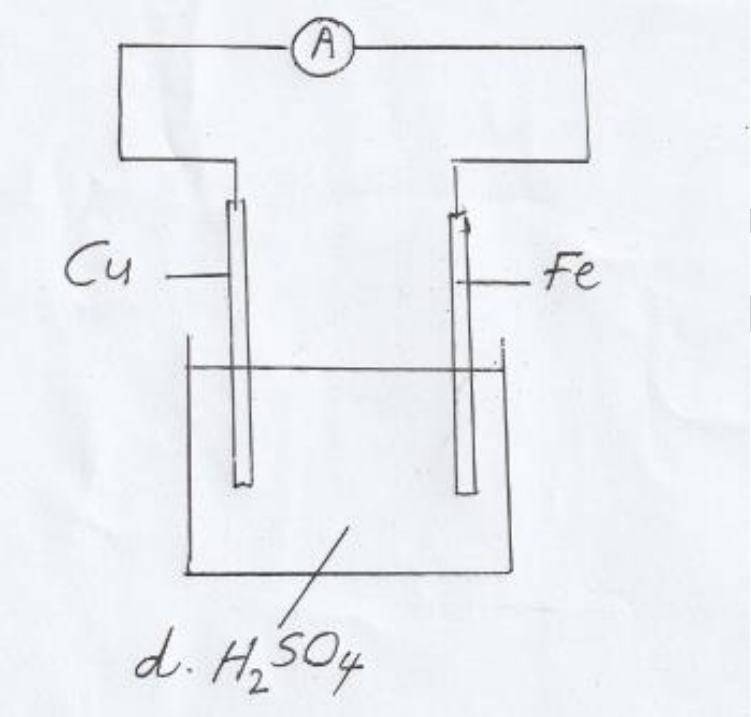
9) Observations.

(i) ……..

(ii) ……..

(iii) ……

Electrochemical cell - 2



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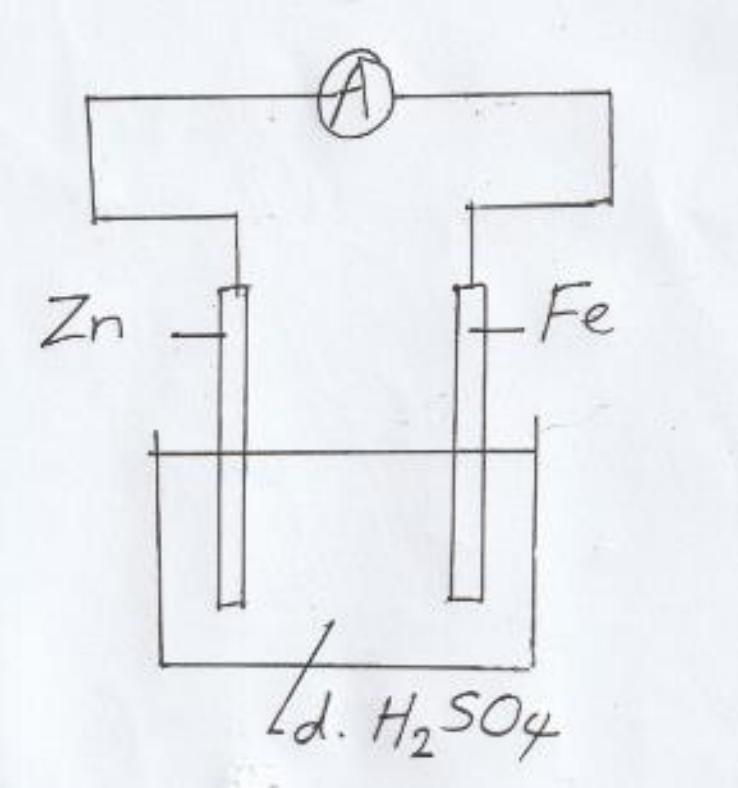
Electrochemical cell - 3



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