Rules when writing chemical formulae

Rule 1 – Write the elements and exchange the valences and write on the lower right side of the elements

Aluminium oxide

Al O

Al_2 O_3

Al_2O_3

Ferric Sulphide

Fe S

Fe_2 S_3

Fe_2S_3

Rule 2 –If the valence is 1, then don't write 1

Sodium Oxide

Na O

Na_2 O_1

Na_2 O

Na_2O

Potassium Chloride

K Cl

K_1 Cl_1

K Cl

KCl

Rule 3 – Before writing a valence to a radical, write a bracket to the radical

- Radicals more made up of more than one element and having a charge.
- I have underlined the radicals in the previous note.

Ammonium Sulphate

NH_4 SO_4

$(NH_4)_2$ $(SO_4)_1$

$(NH_4)_2$ SO_4

$(NH_4)_2$ SO_4

Magnesium Phosphate

Mg PO_4

Mg_3 $(PO_4)_2$

$Mg_3(PO_4)_2$

Rule 4 – Simplify the valences if possible

Plumbic Oxide

Pb O

Pb_2 O_4

Pb_1 O_2

Pb O_2

PbO_2

Magnesium Oxide

Mg O

Mg_2 O_2

Mg_1 O_1

Mg O

MgO

1. Ferric oxide.....
2. Ferrous oxide.....
3. Cuprous oxide.....
4. Cupric oxide.....
5. Copper oxide.....
6. Magnesium chloride.....
7. Sodium chloride.....
8. Calcium hydroxide.....
9. Potassium hydroxide.....
10. Ammonium sulfate.....
11. Ammonium chloride.....
12. Potassium manganite.....
13. Potassium permanganate.....
14. Aluminium chromate.....
15. Aluminium dichromate.....
16. Calcium phosphate.....
17. Aluminium phosphate.....

18. Potassium sulfide……

19. Potassium sulfite……

20. Potassium sulfate……

21. Magnesium nitrite……

22. Magnesium nitrate……

23. Sodium carbonate……

24. Sodium bicarbonate……

25. Sodium sulfite……

26. Sodium bisulfite……

27. Sodium sulfate……

28. Sodium bisulfate……

29. Calcium carbonate……

30. Aluminium carbonate……

31. Aluminium phosphate……

32. Ammonium phosphate……

33. Copper sulfate……

34. Sodium bromide……..

35. Ammonium hydroxide……

36. Ferrous sulphate……..

37. Ferric sulphate……..

38. Magnesium chlorate……

39. Magnesium phosphate……

40. Zinc sulphate……..

41. Ammonium nitrate…….. ……

42. Calcium oxychloride……

43. Zinc sulfide……..

44. Magnesium nitrate……

45. Potassium Bicarbonate……