

Rules when writing chemical formulae

Rule 1 – Write the elements and exchange the valences and write on the lower right side of the elements

Aluminium oxide

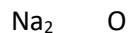
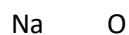


Ferric Sulphide



Rule 2 – If the valence is 1, then don't write 1

Sodium Oxide



Potassium Chloride



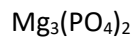
Rule 3 – Before writing a valence to a radical, write a bracket to the radical

- Radicals more made up of more than one element and having a charge.
- I have underlined the radicals in the previous note.

Ammonium Sulphate

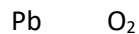


Magnesium Phosphate

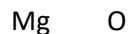
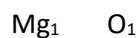
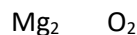
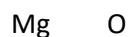


Rule 4 – Simplify the valences if possible

Plumbic Oxide



Magnesium Oxide



1. Ferric oxide.....
2. Ferrous oxide.....
3. Cuprous oxide.....
4. Cupric oxide.....
5. Copper oxide.....
6. Magnesium chloride.....
7. Sodium chloride.....
8. Calcium hydroxide.....
9. Potassium hydroxide.....
10. Ammonium sulfate.....
11. Ammonium chloride.....
12. Potassium manganite.....
13. Potassium permanganate.....
14. Aluminium chromate.....
15. Aluminium dichromate.....
16. Calcium phosphate.....
17. Aluminium phosphate.....

18. Potassium sulfide.....

19. Potassium sulfite.....

20. Potassium sulfate.....

21. Magnesium nitrite.....

22. Magnesium nitrate.....

23. Sodium carbonate.....

24. Sodium bicarbonate.....

25. Sodium sulfite.....

26. Sodium bisulfite.....

27. Sodium sulfate.....

28. Sodium bisulfate.....

29. Calcium carbonate.....

30. Aluminium carbonate.....

31. Aluminium phosphate.....

32. Ammonium phosphate.....

33. Copper sulfate.....

34. Sodium bromide.....

35. Ammonium hydroxide.....

36. Ferrous sulphate.....

37. Ferric sulphate.....

38. Magnesium chlorate.....

39. Magnesium phosphate.....

40. Zinc sulphate.....

41. Ammonium nitrate.....

42. Calcium oxychloride.....

43. Zinc sulfide.....

44. Magnesium nitrate.....

45. Potassium Bicarbonate.....