

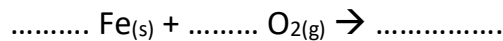
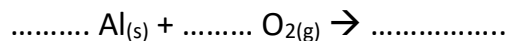
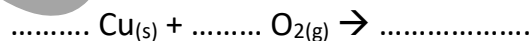
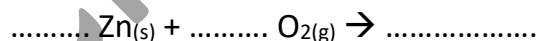
Grade	Institute	Day	Time	Starting Date
10	Channa Asela Institute, Mt. Lavinia	Saturday	02.30pm – 05.30pm	5 th September
10	Shakthi Institute, Colombo 04	Friday	02.30pm – 04.30pm	4 th September
10	Shakthi Institute, Waterfront, Colombo 2	Tuesday	02.30am – 04.30pm	8 th September
11	Channa Asela Institute, Mt. Lavinia	Saturday	10.30am – 01.30pm	5 th September
11	Shakthi Institute, Colombo 04	Monday	02.30pm – 04.30pm	7 th September
11	Shakthi Institute, Waterfront, Colombo 2	Wednesday	02.30pm – 04.30pm	2 nd September
Past Paper	Channa Asela Institute, Mt. Lavinia	Saturday	06.00pm – 08.00pm	5 th September
Past Paper	Shakthi Institute, Colombo 04	Monday	04.45pm – 6.45pm	7 th September
Past Paper	Shakthi Institute, Waterfront, Colombo 2	Wednesday	0.4.45pm – 06.45pm	2 nd September

Reactivity Series

- 1) are arranged in the order according to their of reaction when they react with, and
- 2) K, Na, Ca,, Al, Zn,, Sn, Pb, H,, Hg, Ag,, Au
- 3) The on the left side are reactive and the on the right side are reactive.
- 4) Example Zn is reactive than Fe and Mg is reactive than Zn.

How metals react with O₂ in air

- 1) Since and are very reactive, when they are exposed to, they will react with and produce their metal
- 2) Therefore and will lose their very
- 3) Hence & are stored in oil or oil to prevent them getting exposed to
- 4) from to will react with in air and produce their metal, when they are in



5) Since, &are extremely reactive, they will not react with and produce their metal even if they are in

6) Therefore, & will never lose their

Colours of flames when metals are burnt in air

Li –

Na –

K –

Mg –

Ca –

How metals react with water

1) Since & are very reactive, they react with water and produce their metal and gas.

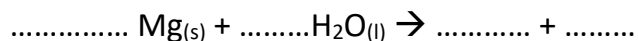
2) Since the density of is than the density of water, floats and darts on water while producing a hissing sound.

3) Sometimes may catch fire.

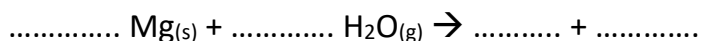


4) does not react with water.

5) reacts with water and produce and gas

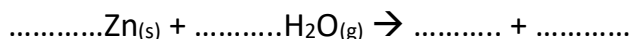
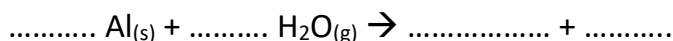


6) also reacts with and produceand gas



7) and do not react with water or water.

8) and react with and produce their metal and gas

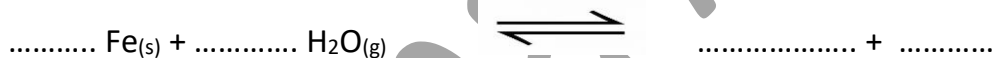


9) does not react with water or water or

10) But reacts with when and produce and gas.

11) This is a reaction. That means reacts with when and produce and gas and also the and gas will react and produce and

12) Therefore reaction means the can become and then again the can become

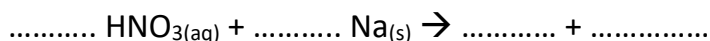
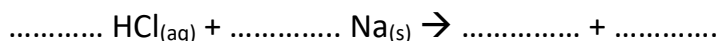
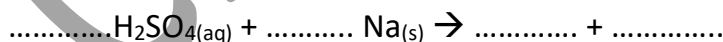


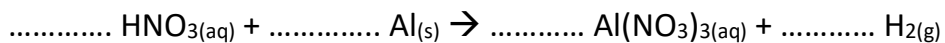
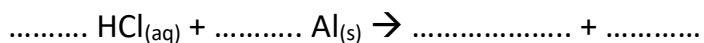
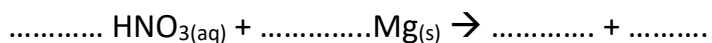
13) Metals such as, & do not react with even when

How metals react with acids

1) The which are reactive than (from to) will react with acids and produce their metal and

2) The which are reactive than will not react with acids (from to))





- $\text{H}_2\text{SO}_{4(\text{aq})} = \dots\dots\dots$
- $\text{HCl}_{(\text{aq})} = \dots\dots\dots$
- $\text{HNO}_{3(\text{aq})} = \dots\dots\dots$
- Since $\dots\dots\dots$, $\dots\dots\dots$, $\dots\dots\dots$ are very reactive, they may even catch $\dots\dots\dots$ and $\dots\dots\dots$ when they react with $\dots\dots\dots$ acids.

Channa Asela